**Chemistry Acids & Bases (Chapter 18 Outline)**

***AS YOU READ, RESPOND TO THE FOLLOWING:***

*Vocabulary*: **Define in your own words**

Acidic solution-

Basic solution-

Conjugate acid-

Conjugate base -

Conjugate acid-base pair-

Acid ionization constant-

Base ionization constant-

pH-

pOH-

Salt-

Titration-

Buffer-

Neutralization reaction-

Respond to the prompts below:

What are the physical and chemical properties of acids and bases? Describe these properties in detail. (4 pts)

How do the Arrhenius, Bronsted- Lowry model, and Lewis models of acids and bases compare? Explain how these models compare and contrast with each other. (3 pts)

How is the strength of an acid or base related to its degree of ionization? Explain. (2 pts)

What is the relationship between the strengths of acids and bases and the values of their ionization constants? Explain. (2 pts)

How are pH and pOH related to the ion product constant for water? Explain. (2 pts)

How are neutralization reactions used in acid-base titrations? Explain. (2 pts)

How do the properties of buffered and unbuffered solutions compare? (1 pt)

***ABOUT THE READING:***

Write three things that you learned about Acids & Bases*:*

***Make sure to write a full sentence.***

*Example: I learned that the equivalence point of a titration reaction is the point where the moles of H+ ions (acid ions) are equal to the moles of OH- ions (base ions).*

*1.*

*2.*

*3.*

***Assigned work****: Chapter 18 assessment, problems #57, 58, 69, 82, 83, 91*